

Calculating Ph Pogil Answer Key

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The pH is given by the equation . Since hydrochloric acid is monoprotic, the concentration of the solution is equal to the concentration of protons. Using this value and the pH equation, we can calculate the pH. Now we can find the pOH. The sum of the pH and the pOH is always 14. The pOH of the solution is 7.8.

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To obtain pH, we could calculate [H 3 O+] = K w /[OH-], and then get pH; or we could calculate pOH from our value of [OH-], and get pH from pH = 14.000 - pOH. Taking the second approach, we have pOH = -log(1.48 x 10-3) = 2.830 Y pH = 14.000 - 2.830 = 11.170 8.
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Making the usual assumptions we can calculate [OH-] and pOH to obtain pH: pOH = 5.239 pH = 14.000 - 5.239 = 8.761 Note that the pH is not 7! For a titration of a weak acid with a strong base titrant, at the equivalence point the pH will always be >7, because we have created a solution of the pure conjugate base in water.

Chem 116 POGIL Worksheet - Week 11 - Solutions Titration
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