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Chapter 17. Name \_\_\_\_\_ STUDY GUIDE. Acids and Bases Use with Text Pages 466–473. Use the words in the box to fill in the blanks in the paragraphs below. An acid is a substance that produces \_\_\_\_\_ in solution. Acids taste \_\_\_\_\_. Some acids can produce painful burns and damage tissues.

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Acids are named by changing the ending of the alkane to -oic acid. i.e. a 4 carbon acid goes from butane to butanoic acid Acid salts are named by changing the acid ending to -oate i.e the 4 carbon acid goes from butanoic acid to sodium butanoate Esters are named by changing the ending of the alkane to -oate and by using alkyl name i.e. the 4 carbon ester below is methyl butanoate(methyl for OCH<sub>3</sub> portion) O Anhydrides are dehydrated esters that replace -oic acid with word anhydride Acyl ...

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bitter taste, slippery feel, turns red litmus paper blue, more.... organic bases found in plants that are often poisonous.

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Carboxyl, amine, hydrogen and variable side group Describe two forces that help shape an amino acid chain? Peptide bond holds chains of amino acids together by linking amine and carboxyl groups; hydrogen bonds help fold proteins into shapes and hold them together. Shape depends on amino acids, polarity and R groups.

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•The most common strong acids are HCl, HBr, HI, HNO<sub>3</sub>, HClO<sub>3</sub>, HClO<sub>4</sub>, and H<sub>2</sub>SO<sub>4</sub>. •Strong acids are strong electrolytes. •All strong acids ionize completely in solution. •Example: Nitric acid ionizes completely in aqueous solution.  $\text{HNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$  •Since H<sup>+</sup> and H<sub>3</sub>O<sup>+</sup> are used interchangeably, we write HNO

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There are 20 amino acids, but only 4 nucleotide bases in DNA. That means it takes at least 3 nucleotides to code for one amino acid, hence a codon (three amino acids). ? The code is read as an RNA code, 5' to 3'.

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