

## Chemistry Review Sheet Unit 7 Answer Key

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are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane, C 2H 6. C 2H 6 2(12.0 g) + 6(1.0 g) = 30.0 g 2. Convert 45 g Se to moles of Se. ? mol Se = 45 g Se ! 1 mol Se 79.0 g Se = 0.57 mol Se 3.

~~Review Sheet Unit 7 Name Vocabulary Fill in the blanks~~

are representative of each type of problem worked in this unit. For more practice on a given type of problem, refer back to Note Taking Guides, worksheets, and quizzes. 1. Calculate the molar mass for ethane, C 2H 6. 2. Convert 45 g Se to moles of Se. 3. Convert 4.77 x 1024 molecules of SO 2 to grams.

~~Review Sheet Unit 7 Name Vocabulary Fill in the blanks~~

UNIT 7 REVIEW SHEET - EQUILIBRIUM [] EQUILIBRIUM is the point at which the rate of the forward reaction is equal to the rate of the reverse reaction in a reversible process. o Reaction is dynamic – still occurring – even though to the eye, it looks like nothing is changing at equilibrium. o Observable/measurable quantities (concentration, pressure, color) are constant at equilibrium. o The amounts of products and reactants are typically not equal at equilibrium!

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CHEMISTRY CP Name: \_\_\_\_ Per: \_\_\_\_ TEST: \_\_\_\_ Unit 7 Review Sheet: Thermochemistry KEY Potential Energy Diagrams: Choose the appropriate letter. 1. What is the ΔH of the forward reaction? f 2. What is the activation energy of the forward reaction? b 3. What is the H of the reverse reaction? f 4.

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Review Sheet: Unit 5 Name \_\_\_\_ CHEMISTRY: A Study of Matter © 2004, GPB 5.24a KEY I. For each of the following statements, write “I” for ionic, “C” for covalent

~~CHEMISTRY–A Study of Matter~~

Review Sheet for Chemistry Unit 5 December 11, 2011 31. Balance the following equations by putting the proper coefficients wherever they are needed. a) Sr + O2 SrO b) Al + H 2O Al(OH)3 + H 2 c) Al + O 2 Al2O3 d) C5H12 + O 2 CO 2 + H 2O e) Na + H 2O H2 + NaOH f) NaClO 3 NaCl + O2 g ...

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