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The infra-red spectra of H, I and J are given below, but it does indicate which is - which. Identify the three compounds H, I and J, using the infra-red spectra below,

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and decide which spectrum belongs to which compound. (iii) Compound E, which is a branched chain haloalkane, was found to have the composition by mass of 39.8% C, 7.3% H, 2)

## IR TASK 1 - Weebly

1) 2) 3) 4) 5) 6) bromoethane — ¥ ethyl ethanoate cold, aq NaOH nucleophilic substitution propene 2-methylpropylamine chloroethane ethanoic acid

## ORGANIC SYNTHESIS PROBLEMS

2-bromo-2-methylbutane 4 5-40 (CH<sub>3</sub> CH<sub>2</sub>), 5-40 (CH<sub>3</sub> CH<sub>2</sub>), 10-70 (CBr), 5-40 ((CH<sub>3</sub>)<sub>2</sub> CBr) methylpropene ) 3 90-150 (CH<sub>2</sub>), 90 -150 (C<sub>5</sub> 40 H 3 2) propene 3 ) 90-150 (CH<sub>2</sub>), 90 -150 (C<sub>5</sub> 40 H 3 2-chloropropane ), 2 5-40 (CH<sub>3</sub> 10-70 (CHCl) propanone ), 2 20-50 (CH<sub>3</sub> 190-220 (CO) methylamine 1 25-60 (CH<sub>3</sub>)

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2,2-dimethylamine 2-methylbut-1-ene

3-methylbutan-1-ol 4-hydroxybutanone

3-methylbutanoic acid butyl methanoate

ethanenitrile propanal Decide which

spectrum belongs to which compound and draw the molecule next to the spectrum.

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ProEpi

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Propene reacts with HBr to form H. H reacts with sodium hydroxide to form I, and I reacts with warm acidified potassium dichromate (VI) to form J. The infra-red spectra of H, I and J are given below, but it does indicate which is - which. Identify the three compounds H, I and J, using the infra-

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H Nmr Spectroscopy Answers

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1 use your knowledge of ionic equations to

give the molar ratio in which the following

acids react with bases complete the table

to show your answers acid formula of acid

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the rate equation for each of the following  
reactions a a reacts with b to form c a b c  
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compounds are shown. Complete the table  
to match the spectra to the compounds.  
Identify any key signals you used to  
identify each compound. You may not be  
able to decide between two of the  
compounds. Structure Spectrum Bond(s)  
Wavenumber range (cm<sup>-1</sup>) Structure  
Spectrum Bond(s) Wavenumber range (cm  
l) butanoic ...

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compounds are shown ...

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rate =  $k[A][B]^2$   $k = \frac{\text{rate}}{[A][B]^2} = \frac{16}{2.0 \times 2.0^2} = 2.0 \text{ mol}^{-2} \text{ dm}^6 \text{ s}^{-1}$   
a) rate =  $k[E]^2$   $k = \frac{\text{rate}}{[E]^2} = \frac{0.80}{0.20^2} = 20 \text{ s}^{-1}$   
b) rate =  $k[E]^4$   $k = \frac{\text{rate}}{[E]^4} = \frac{0.20}{0.20^4} = 0.20 \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$   
c) rate =  $k[P][Q]$   $k = \frac{\text{rate}}{[P][Q]} = \frac{0.00200}{0.1 \times 0.1} = 0.20 \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$   
d) rate =  $k[H^+][CH_3COCH_3]$   $k = \frac{\text{rate}}{[H^+][CH_3COCH_3]} = \frac{9 \times 10^{-5}}{0.5 \times 6.0} = 3.0 \times 10^{-5} \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$

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