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mol) = 474 kJ / mol. Knowing that $\Delta G = -nFE^{\circ}$ and $n = 4$, we calculate the potential is -1.23 V.

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Thermodynamics – problems and solutions. The first law of thermodynamics. 1. Based on graph P-V below, what is the ratio of the work done by the gas in the process I, to the work done by the gas in the process II? Known : Process 1 : Pressure (P) = 20 N/m 2. Initial volume (V_1) = 10 liter = 10 dm³ = 10 x

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Answers For Thermodynamics Problems
Answer for Problem # 1 Since the
containers are insulated, no heat transfer
occurs between the gas and the external

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environment, and since the gas expands freely into container B there is no resistance "pushing" against it, which means no work is done on the gas as it expands.

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Processes (Ideal Gas) A steady flow compressor handles $113.3 \text{ m}^3/\text{min}$ of nitrogen ($M = 28$; $k = 1.399$) measured at intake where $P_1 = 97 \text{ KPa}$ and $T_1 = 27 \text{ C}$. Discharge is at 311 KPa . The changes in KE and PE are negligible. For each of the following

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The first law of thermodynamics –

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problems and solutions. 1. 3000 J of heat is added to a system and 2500 J of work is done by the system. What is the change in internal energy of the system? Known : Heat (Q) = +3000 Joule. Work (W) = +2500 Joule . Wanted: the change in internal energy of the system. Solution : The equation of the first law of

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The entropy, S , increases because there are more moles of gaseous products. (e) $\text{CH}_3\text{COOH} (l) \rightarrow \text{CH}_3\text{COOH} (s)$ $\Delta S^\circ < 0$ (decrease in entropy) since the reaction goes from a liquid to a solid. (f) $\text{N}_2 (g) + \text{O}_2 (g) \rightarrow 2 \text{NO} (g)$ $\Delta S^\circ \sim 0$ (little change in entropy) since the reaction goes from 2

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moles of gases to 2 moles of gases.

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What is the change in internal energy of the system. Known : Heat (Q) = +3000 Joule Work (W) = +2500 Joule Wanted : the change in internal energy of the system
Solution : Equation of the first law of thermodynamics $\Delta U = Q - W$ The sign

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With Solutions : Q is positive if the heat added to the system W is positive if work is done by the system Q is negative if heat leaves the system W is negative if work is done on the system The change in internal energy of the system : $\Delta U = 3000 - 2500$ $\Delta U = 500 \dots$

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Properties 1. If an object has a weight of

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10 lbf on the moon, what would the same object weigh on Jupiter? Jupiter...

~~Thermodynamic Properties~~

Physics problems: thermodynamics. Part 1
Problem 1. A rapidly spinning paddle wheel raises the temperature of 200mL of water from 21 degrees Celsius to 25

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degrees. How much a) work is done and b) heat is transferred in this process? Solution . Problem 2. The temperature of a body is increased from -173 C to 357 C .

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$\rho gh = P_{\text{pipe}} - P_{\text{atm}}$ and therefore, $h =$

$$(P_{\text{pipe}} - P_{\text{atm}}) / \rho g = (135000\text{ Pa} - 92000$$

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Pa) / (1000 kg/m³ x 9.81 m/s²) = 4.4 m.

School of Engineering, University of
Edinburgh Engineering Thermodynamics
2 and Thermodynamics (Chemical) 2.

Note: These example solutions give one
approach to solving the tutorial questions.

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